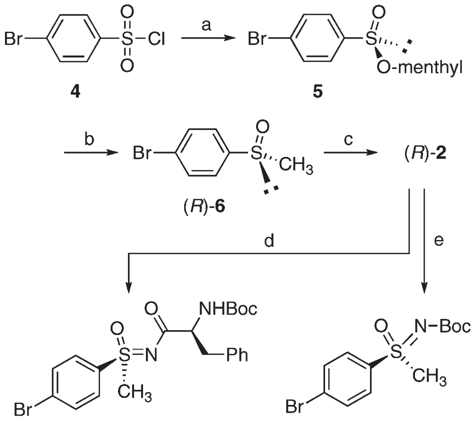
JOC *Additions and Corrections*

Correction to Synthesis and Palladium-Catalyzed Coupling Reactions of Enantiopure *p*-Bromophenyl Methyl Sulfoximine [*J. Org. Chem.* **2005,** *70*, 2346. DOI: 10.1021/jo047940c]. Gae Young Cho, Hiroaki Okamura, and Carsten Bolm*

Page 2347. Due to a stereochemical misassignment of compound **5**, the absolute configuration of product **6** is incorrect. Sulfoximine **2** as well as all subsequent molecules derived thereof should have the *R*-configuration at sulfur. A correct Scheme 1 is shown below. We thank M. Frings and Dr. I Atodiresei (both RWTH Aachen University) for alerting us to this issue.

SCHEME 1



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Why Are Esters and Amides Weaker Carbon Acids than Ketones and Acid Fluorides? Contributions by Resonance and Inductive Effects [*J. Org. Chem.* **2009**, *74*, 7245. DOI: 10.1021/jo901225t]. Alexandra Fersner, Joel M. Karty,* and Yirong Mo*

Page 7251. We suggested that the large difference between the resonance contribution we calculated for acetone (34.3 kcal/mol) and the one Silva calculated for acetaldehyde (41.9 kcal/mol)[1] “might be due to geometrical constraints imposed on the perpendicular vinylogue structures. It is not clear that Silva carried out the same optimizations. If he did not, then the resonance contribution he reports will be artificially high”. After communicating with Silva, we now believe that this is not the reason for the difference, as his original paper does indeed mention that such optimizations were performed. Instead, the difference is largely due to the different methods employed. We used an extrapolation methodology on the $n = 1-4$ vinylogues to arrive at the resonance contribution for the $n = 0$ species, and worked with thermally corrected enthalpies. Silva derived his numbers from direct calculations on the strained $n = 0$ species, working with bottom-of-the-well energies. Silva has graciously applied his methodology to acetone, and the results suggest a resonance contribution of 41.4 kcal/mol, nearly identical to what he calculated for acetaldehyde.[2] Inclusion of zero-point vibrational energy corrections reduces this number to 39.1 kcal/mol, bringing the numbers from the two methodologies in somewhat better agreement.

(1) Silva, P. J. *J. Org. Chem.* **2009**, *74*, 914−916.
(2) Silva, P. J. Personal communication, 2009.

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